

For each of the questions only **ONE** of the lettered responses (A-D) is correct.

Select the correct response in each case and mark its code letter by connecting the dots as illustrated on the answer sheet.

1 Given the bond energies:

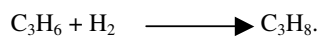
$$\text{C} - \text{C} \quad 348 \text{ kJ mol}^{-1}$$

$$\text{C} - \text{H} \quad 412 \text{ kJ mol}^{-1}$$

$$\text{C} = \text{C} \quad 612 \text{ kJ mol}^{-1}$$

$$\text{H} - \text{H} \quad 436 \text{ kJ mol}^{-1}$$

calculate the enthalpy change for the reaction



A -248 kJ mol^{-1}

B $+248 \text{ kJ mol}^{-1}$

C -124 kJ mol^{-1}

D $+124 \text{ kJ mol}^{-1}$